Additions and Corrections

^{13}C NMR Studies on Arsenic(III) and Antimony(III) Dihydroxydicarboxylate Complexes [*J. Am. Chem. Soc.* **1980**, *102*, 5712]. Dragoslav Marcovich and Robert E. Tapscott.*

Page 5713, the sixth and seventh sentences of the first paragraph should be changed to read: The first crop of crystals obtained, upon evaporation of the acetone, was racemic material while the last crop obtained was essentially pure (+)-dimethyltartaric acid. Workup of the filtrate, after removal of the quinine salt of (+)-dimethyltartrate, in a similar manner gave (−)-dimethyltartaric acid.

Classical Electrical Contributions to Solvent Polarity Scales [*J. Am. Chem. Soc.* **1981**, *103*, 6036]. S. Ehrenson.

Page 6042: Equation A7 should read:

$$W_l(w) = (-1)^{l+1}e^{-w}\sum_{k=0}^{l}\frac{(l+k)!(2w)^{l-k}}{k!(l-k)!} + e^{w}\sum_{k=0}^{l}\frac{(-1)^k(l+k)!(2w)^{l-k}}{k!(l-k)!}$$

Equation A8 should read:

$$V_l(w) = w^{2l}\,\mathrm{d}^1\,(\sinh w/w)/(w\,\mathrm{d}w)^l$$

The Automerization of $C_{11}H_{11}$ Chlorides and the Stability of Their Cations [*J. Am. Chem. Soc.* **1981**, *103*, 6530]. M. J. Goldstein,* S. Tomoda, E. J. Pressman, and J. A. Dodd.

The sixth letter of the Greek alphabet (ζ, zeta) is correctly indicated at the sixth carbon atom of **10**. It was inadvertently replaced by the fourteenth letter (ξ, xi) in the text that follows and in Table II.

Thermal Rearrangement of an Allenic Diazoalkane and Intermolecular Capture of a Diazoethene by a Cyclopropene To Give a Common Dihydropyridazine Product [*J. Am. Chem. Soc.* **1981**, *103*, 7011]. Paul M. Lahti and Jerome A. Berson.*

Page 7011, Table I: The superscript *d* should refer to the third (not the second) structural formula in the column headed "allenes", whereas the label 32^a, 20^b should refer to the second formula, not the third.

Base Catalysis in a Photochemical Smiles Rearrangement. A Case of General Base Catalysis of a Photoreaction [*J. Am. Chem. Soc.* **1981**, *103*, 7669]. Gene G. Wubbels* and Daniel W. Celander.

Page 7670: Equations 1, 2, and 4 are incorrect due to the presence of the partitioning factor, *f*, in the denominator of eq 1. The proper equations are:

$$\Phi = \Phi_{\mathrm{ISC}}\left(\frac{k_2}{k_1+k_2}\right)\left(\frac{k_4+fk_5[\mathrm{B}]}{k_3+k_4+k_5[\mathrm{B}]}\right) \qquad (1)$$

$$\frac{1}{\Phi} = \frac{1}{\Phi_{\mathrm{ISC}}}\left(1+\frac{k_1}{k_2}\right)\left(\frac{k_3+k_4+k_5[\mathrm{B}]}{k_4+fk_5[\mathrm{B}]}\right) \qquad (2)$$

$$\frac{1}{\Phi} = \frac{1}{\Phi_{\mathrm{ISC}}}\left(1+\frac{k_1}{k_2}\right)\left(\frac{1}{f}+\frac{k_3+k_4}{fk_5[\mathrm{B}]}\right) \qquad (4)$$

This revision changes slightly the values of the calculated rate constants; the revised values are the following: $k_3 = 7.1 \times 10^7$ s^{-1}, $k_4 = 1.8 \times 10^7$ s^{-1}, $k_5^{\mathrm{OAc}} = 2.0 \times 10^7$ M^{-1} s^{-1}, and $k_5^{\mathrm{HPO_4}} = 4.6 \times 10^8$ M^{-1} s^{-1}.

Studies on Spin-Trapped Radicals in γ-Irradiated Aqueous Solutions of L-Alanylglycine and L-Alanyl-L-alanine by High-Performance Liquid Chromatography and ESR Spectroscopy [*J. Am. Chem. Soc.* **1982**, *104*, 830]. Fumio Moriya,* Keisuke Makino, Nobuhiro Suzuki, Souji Rokushika, and Hiroyuki Hatano.

Page 833, caption to Figure 7, line 4: Na_2HPO_4–NaOH buffer, pH 11.5 should be replaced by H_3BO_3·NaCl–NaOH buffer, pH 9.0.

Page 833, right column, line 1: L-**Alanyl**-D-**alanine** should be replaced by L-**Alanyl**-L-**alanine**.

Nuclear Magnetic Resonance Investigation of the Spontaneous Decarboxylation of 2-Oxalopropionic Acid. 2. Species in Solution [*J. Am. Chem. Soc.* **1981**, *103*, 7609]. G. Kubala and A. E. Martell.*

Page 7614: Equation 7 should read:

$$\alpha_{\mathrm{HA}} = (10^{pK_1 - pD} + 1 + 10^{pD - pK_2})^{-1}$$

*Book Reviews**

Fourier Transform Infrared Spectroscopy. Applications to Chemical Systems. Volume 1. Edited by J. R. Ferraro and L. J. Basile (Argonne National Laboratory). Academic Press, New York. 1978. viii + 311 pp. $34.00.

The present volume on the applications of Fourier-transform infrared spectroscopy to chemical systems reflects a renaissance in the use of infrared spectroscopy that has occurred over the past decade. Due to major technological advances brought about by computer-dedicated Fourier-transform spectroscopy, infrared absorption is making deep inroads into research areas that have been either untouched, previously abandoned, or dominated by laser-Raman scattering. It is fitting then that this new methodology be given the exposure that it deserves in a book of high quality and professional character.

The first chapter by Green and Reedy on matrix-isolation FTIR spectroscopy provides an excellent introductory chapter. Both the basic principles and advantages of FTIR spectroscopy as well as the rudiments of the matrix-isolation technique are amply described. Furthermore, the matrix-isolation method is well-suited to demonstrating the advantages of the FTIR approach to obtain high-quality, high-resolution spectra of molecular samples that pose severe difficulties with more conventional approaches.

The ensuing chapters take up other topics of fundamental interest and importance and should appeal to those engaged in research, as well as in analytical applications. In the second chapter D'Esposito and Koenig describe applications of FTIR to synthetic and biological polymers. A major focus of this chapter is to demonstrate the subtractive accuracy of FTIR to obtain individual component spectra from mixtures and to eliminate unwanted solvent absorption bands. The latter technique shows particular promise for the study of biological samples in aqueous media,

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*Unsigned book reviews are by the Book Review Editor.